

# Electrochemistry of Chemically Trapped Dimeric and Monomeric Recombinant Horseradish Peroxidase

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## Abstract

Native horseradish peroxidase (nHRP) exists in the aggregated form in concentrated water solutions as shown by dynamic light scattering (DLS). This is in contrast to recombinant horseradish peroxidase (reCHRP) which mainly exists as a dimer. The native enzyme aggregates could be broken into the particles of nm-size only under the conditions of high ionic strength (0.5-1 M NaCl). Chemical cross-linking of reCHRP with glutaraldehyde in water solutions yields 40% of the dimer. The chemically trapped dimeric and monomeric forms of reCHRP were separated by gel-filtration, their substrate specificity towards a number of organic substrates compared. Parameters of direct and mediated electron transfer on graphite electrodes catalyzed by both preparations were analyzed. The difference in behavior of the monomeric and dimeric enzyme forms observed in electrochemical experiments was interpreted as a result of a "double" coverage of the electrode surface with the molecules of cross-linked dimeric enzyme, in contrast to both modified monomeric and original, unmodified reCHRP providing "monolayer" coverage. In addition to the stabilization effects achieved due to enzyme surface modification with glutaraldehyde, the "double" coverage doubles the enzyme activity per surface unit.

## Keywords

*Dimeric Recombinant Peroxidase; Light Scattering; Direct Electron Transfer; Electrochemical Kinetics; Monolayer; Bilayer; Graphite Electrode*

## Introduction

Horseradish peroxidase HRP (EC 1.11.1.7) is one of the enzymes most widely used for applied purposes and in biosensors in particular. The enzyme is usually immobilized on electrode surfaces by adsorption from concentrated solutions (10 mg/mL). Despite a lot of enzyme has been utilized, immobilization procedures yield rather low activity, commonly explained by enzyme aggregation upon interaction with the electrode surface. Native HRP needs covalent immobilization on electron-conducting polymer films to catalyze direct electron transfer reactions (Ran et al, 2011), whereas reCHRP lacking oligosaccharides (Gazaryan et al., 1994; Gazaryan et al., 1997) catalyzes direct electron transfer (DET) reactions (Gorton et al, 1999, Lindgren et al., 1999, Shipovskov et al., 2004, Ferapontova et al., 2005, Kartashov et al., 2010). The principal improvement in bioelectrochemical properties of reCHRP is explained by the absence of oligosaccharide chains screening the active site from direct contact with the electrode material.

The oligomeric state of both nHRP and reCHRP in water solutions has never been the focus of research paper. It was found earlier that reCHRP exhibited a tendency to form dimeric structures in the system of reversed micelles (Gazaryan et al., 1997). RecHRP upon

binding its inhibitor, benzohydroxamic acid, forms a kind of a dimer in the crystal cell (Henriksen et al., 1998). Both enzymes, nHRP and recHRP, have never been studied by dynamic light scattering (DLS) techniques in water solutions. This paper presents a comprehensive study answering the question on different oligomeric state of both enzymes in water solutions and includes DLS of native and recHRP, chemical trapping of dimeric and monomeric recHRP forms and their comparison in terms of substrate specificity and electrochemical behavior on graphite electrodes. The discovery of heavily aggregated native enzyme in water solutions may explain the advantages of using recHRP for the purposes of construction of biosensors based on DET.

## Materials and Methods

### Chemicals and Reagents

Guaiacol was from Fluka (Buchs, Switzerland). 2, 2'-azino-bis(3-ethylbenzthiazoline-6-sulfonic acid) (ABTS), *o*-phenylenediamine (*o*-PD), *o*-dianisidine (*o*-DA), pyrogallol, glutaraldehyde, sodium bihydride, Tris were from Sigma (St. Louis, MO). *p*-Cresol, hydrogen peroxide (30% H<sub>2</sub>O<sub>2</sub>), buffers and salts were from Merck (Darmstadt, Germany). All aqueous solutions were prepared using water purified with a Milli-Q system (Millipore, Bedford).

### Enzyme Preparation

Native horseradish peroxidase (HRP), RZ=3.0, was purchased from Biozyme (UK). Wild-type recombinant HRP was produced with a scaled-up protocol based on the method developed earlier (Gazaryan et al., 1994). The refolding of 1 g of inactive protein solubilized from inclusion bodies (obtained from 60 g of BL21(DE3)/pSA261 biomass) was performed in 10 L refolding medium containing 1 mM DTT, 0.5 mM oxidized glutathione, 5 mM CaCl<sub>2</sub>, 5  $\mu$ M hemin and 2 M urea in 50 mM Tris-HCl buffer, pH 8.0, for 4 days at 4°C, in accord with the solubilization protocol described in (Gazaryan et al., 1994). The refolding medium was concentrated up to 20-fold and the concentration of urea was reduced to 0.2 mM using a flow-through membrane concentrator. Then, the pH value was adjusted to 4.5 to precipitate the unfolded, aggregated protein. The supernatant solution obtained after centrifugation was purified by gel-filtration on a Sephacryl S-200 column (Pharmacia Biotech, Germany). The procedure yielded a homogeneous preparation of wild-type recombinant HRP (ca. 250 mg, RZ> 3.0). This

preparation was used in this work and others published before (Gazaryan et al., 1997; Gazaryan et al., 1999; Lindgren et al., 1999, Abad et al., 2002, Shipovskov et al., 2004, Ferapontova et al., 2005).

### Samples for Light Scattering

50 mM Na-acetate buffer, pH 5.0, K-phosphate and Tris-HCl buffers, pH 8.0, were used to prepare the 0.1-0.5 mg/ml enzyme solutions. The solutions were filtered through 0.2  $\mu$ m Sartorius Minisart filter into a Ø10 mm quartz cuvette.

### Dynamic Light Scattering

Dynamic Light Scattering DLS measurements were carried out on a Malvern 4700c System equipped with a 128-channel digital correlator. An argon ion laser operated at a 488 nm wavelength was used as the light source. The scattering angle was kept constant at 90°. The quality of measurements was tested over the signal-to-noise ratio, the range of the correlation function, and the difference between the measured and calculated baselines. The particle size distributions were calculated by means of the optimized Laplace inversion methods (CONTIN), whereas a contribution of different modes of molecular motion into the integral size distribution and characteristic particle sizes were determined using cumulant analysis for different portions of the correlation function. Both software programs were supplied by the manufacturer. All measurements were performed at (25.0±0.1) °C.

### Cross-Linking

The cross-linking procedure was performed using a 20-fold excess of glutaraldehyde with respect to the enzyme (10<sup>-4</sup> M enzyme, 2×10<sup>-3</sup> M glutaraldehyde) in 50 mM K-phosphate buffer, pH 8.0, by incubation (3h) at room temperature. Then the mixture was treated with 25  $\mu$ L of NaBH<sub>4</sub> solution (4.8 mg/ml). The products of cross-linking were analyzed by SDS-PAGE.

### Separation of Chemically Trapped Dimeric and Monomeric RecHRP

Separation of chemically trapped dimeric and monomeric recHRP was performed by FPLC gel-filtration on a Superose-12 column (Pharmacia, Uppsala, Sweden) in 0.01 K-phosphate buffer, pH 7.5, containing 0.15 M NaCl.

### Specific Activity Measurements

Specific activity towards ABTS, *o*-DA, *o*-PD, guaiacol,

and pyrogallol was measured on a Shimadzu UV-1602 spectrophotometer (Japan) in accordance with the following protocols. ABTS (0.5 mM) oxidation was carried out in 0.1 M Na-citrate/phosphate buffer, pH 4.8, at 414 nm, using an extinction coefficient of  $31.1 \text{ mM}^{-1}\text{cm}^{-1}$  (Arnao et al., 1990). *o*-Phenylenediamine (2.2 mM) oxidation was studied in 0.1 M Na-citrate/phosphate buffer, pH 5.5, at 445 nm, using an extinction coefficient of  $11.1 \text{ mM}^{-1}\text{cm}^{-1}$  (Porstmann et al., 1985). *o*-Dianisidine (0.1 mM) oxidation was studied in 0.1 M Na-citrate/phosphate buffer, pH 5.5, at 460 nm using an extinction coefficient of  $30 \text{ mM}^{-1}\text{cm}^{-1}$  (Ugarova et al., 1979). Guaiacol (3 mM) oxidation was studied in 0.05 M potassium-phosphate buffer, pH 7.0, at 470 nm using an extinction coefficient of  $5.7 \text{ mM}^{-1}\text{cm}^{-1}$  (Hosoya et al., 1967). Pyrogallol (1 mM) oxidation was studied in 0.025 M Na-citrate/phosphate buffer, pH 6.5, at 420 nm using an extinction coefficient of  $4.4 \text{ mM}^{-1}\text{cm}^{-1}$  (Cioprada et al., 1978). The concentration of hydrogen peroxide used was kept equal to 1 mM. The extinction coefficient for hydrogen peroxide was taken as  $72.7 \text{ M}^{-1}\text{cm}^{-1}$  at 230 nm (George, 1953). The concentration of the recombinant enzymes was equal to  $10^{-9}$  M. Recombinant HRP concentration was calculated from the Soret band absorbance ( $\epsilon_{403} = 100,000 \text{ M}^{-1}\text{cm}^{-1}$  (Smith et al., 1990)).

### Steady-State Kinetics of ABTS Oxidation

Steady-state kinetics was measured on a plate reader (Molecular Devices, USA). Rate constants for ABTS oxidation were calculated from double-reciprocal plots by varying concentrations of both substrates, i.e. hydrogen peroxide and ABTS, in the range of 0.008-0.4 mM and 0.1-0.5 mM, respectively. Enzyme concentrations were equal to 0.26 nM.

### Computer Modeling

Dimer modeling and structure optimization were performed as previously described (Gazaryan et al., 1997) using the crystal structure coordinates of the recombinant HRP (PDB code 1ATJ).

### Electrode Preparation

Peroxidase electrodes were prepared using rods of solid spectroscopic graphite (SGL Carbon, Werke Ringsdorff, Bonn, Germany, type RW001, 3.05 mm diameter). The graphite rods were first polished on wet fine-structured emery paper (grit size: P1200) and then additionally polished on paper to obtain a mirror-like surface. The electrode rods were carefully rinsed with

deionized water and allowed to dry at room temperature. 8  $\mu\text{L}$  aliquot of a peroxidase solution in 20 mM phosphate buffer, pH 7, was added to each of the polished ends of the graphite rods and the adsorption was allowed to proceed for 1 h at  $4^\circ\text{C}$ . The enzyme electrodes were then thoroughly rinsed with 0.1 M sodium phosphate buffer, pH 7.0, and if not immediately used, they were stored in the same buffer at  $4^\circ\text{C}$ . Weakly adsorbed peroxidase was desorbed before measurements, by rotating the electrode in buffer for at least 30 min.

### Amperometric Measurements

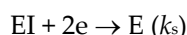
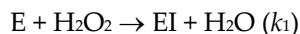
The peroxidase modified electrode was fitted into a rotating disk electrode holder (RDE; EG&G Model 636, Princeton Applied Research, Princeton, NJ), which was placed in a three-electrode cell with an Ag|AgCl (3 M KCl) reference electrode (EG&G) and a platinum wire auxiliary electrode. The electrodes were connected to a BAS 100W potentiostat (Bioanalytical Systems, West Lafayette, IN). All measurements were performed at an applied potential of 0 mV versus Ag|AgCl (Ruzgas et al., 1995; Lindgren et al., 1998) with an electrolyte of 0.1 M phosphate buffer, pH 7.0. The current was registered using various rotation speeds (200–1000 rpm) and  $\text{H}_2\text{O}_2$  concentrations. For mediated electron transfer, a *p*-cresol concentration of 60  $\mu\text{M}$  was used. Prior to experiments, the solution was bubbled with argon for 15 min. Argon was passed over the solution during the experiments.

### Calculation of Kinetic Parameters from Electrochemical Measurements

The reduction current for  $\text{H}_2\text{O}_2$  at a peroxidase-modified electrode is limited by mass-transfer of  $\text{H}_2\text{O}_2$  to the electrode and by the kinetics of the enzymatic reaction. The measured current,  $I$ , is a combination of both the mass-transfer limited current,  $I_{\text{lim}}$  and the kinetically limited current,  $I_{\text{kin}}$ , according to the following equation:  $1/I = 1/I_{\text{lim}} + 1/I_{\text{kin}}$ , which holds for the rotating disk electrode used in this study.  $I_{\text{kin}}$  is calculated by extrapolating the dependence of the registered current ( $I$ ) on the rotation speed,  $\omega$ , in Koutecky–Levich plots ( $I^{-1}$  vs.  $\omega^{-1/2}$ ): the intercept of the Koutecky–Levich plot is equal to the inverse of  $I_{\text{kin}}$ .

In the absence of an electron donor other than the electrode, the process is usually referred to as direct electron transfer (DET) (Gorton et al., 1999). In this case (see Fig.1A) the peroxidase catalytic cycle can be presented as follows:

## SCHEME 1



where E and EI are ferric and 2e-oxidized states of enzyme, respectively.



FIG. 1 SCHEMATIC REPRESENTATION OF DIRECT (A) AND MEDIATED (B) ELECTRON TRANSFER IN THE CASE OF PEROXIDASE-GRAPHITE ELECTRODES ILLUSTRATING THE STATEMENT THAT ONLY A PART OF THE IMMOBILIZED PEROXIDASE MOLECULES CAN ACCEPT ELECTRONS DIRECTLY FROM THE ELECTRODE, IN CONTRAST TO DIFFUSING SUBSTRATE RADICALS ( $S^*$ )

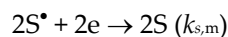
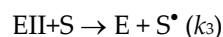
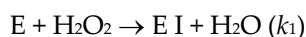
The dependence of the kinetic current on  $H_2O_2$  in this case is described by the following equation (Gorton et al., 1999; Lindgren et al., 1998; Ruzgas et al., 1995):

$$1/I_{kin} = (1/k_1c^* + 1/k_s)/nFE_{DET},$$

where  $n$  is the number of electrons transferred per  $H_2O_2$  molecule ( $n = 2$ ),  $c^*$  is the bulk concentration of  $H_2O_2$ ,  $E_{DET}$  is the amount of adsorbed enzyme active in direct ET and  $F$  is the Faraday constant ( $96\,484\text{ C mol}^{-1}$ ). The dependence of  $I_{kin}$  vs.  $[H_2O_2]$  in double reciprocal plots allows the kinetic constants  $k_1$  and  $k_s$  to be evaluated from the slope and intercept, respectively (Lindgren et al., 1998; Ruzgas et al., 1995).

In the presence of a peroxidase electron donor in a saturating concentration sufficient to engage all peroxidase molecules into mediated ET as shown in Fig. 1B, the peroxidase reaction cycle can be presented as:

## SCHEME 2



where E, EI and EII are ferric peroxidase and its oxidized compounds I and II, respectively, S and  $S^*$  are one-electron peroxidase electron donor and its oxidized form.

Among all rate constants controlling the reduction steps,  $k_3$  is usually considered to be rate-limiting. In this case  $I_{kin}$  can be expressed as follows:

$$1/I_{kin} = (1/k_1c^* + 1/k_3[S])/2m_1FE,$$

where  $m_1$  is the number of electrons transferred per donor molecule ( $m_1=1$ ),  $E$  is the total amount of active enzyme on the electrode surface,  $k_3$  is the rate-limiting

constant for the enzyme reduction by the peroxidase substrate (Gorton et al., 1999; Lindgren et al., 1998). By presenting the experimental data in double-reciprocal plots  $1/I_{kin}$  vs.  $1/[H_2O_2]$ , the kinetic constants  $k_1$  and  $k_3$  can be evaluated. The basic assumption is that in the presence of saturating concentrations of a soluble electron donor, acting as an ET mediator, all peroxidase molecules work in mediated ET, whereas in its absence only a portion of the molecules is active in direct ET. The ratio of the slopes obtained in the presence and absence of a saturating concentration of an electron donor (*p*-cresol) is the ratio between the enzyme active in direct ET ( $E_{DET}$ ) and the total amount of active enzyme ( $E$ ) (Gorton et al., 1999; Lindgren et al., 1998). It should be emphasized that the ratio between the enzyme active in direct ET and the total enzyme is not dependent on the amount of enzyme on the electrode surface and thus, various peroxidase forms can easily be compared on this basis.

The rate constants ( $k_s$ ,  $k_1$  and  $k_3$ ) can be calculated, if the enzyme concentration on the electrode is known. The latter is a serious and still unsolved problem for peroxidases adsorbed on graphite surfaces (Lindgren et al., 1998). Peroxidases are not expected to adsorb on solids with a higher coverage than a monolayer (Andrade et al., 1986). In our calculations, we therefore assume for monolayer coverage of adsorbed peroxidase on the electrode surface (with a roughness factor of 5 (Jaegfeldt et al., 1983)) to be equal to  $40\text{ pmol cm}^{-2}$  (Paddock et al., 1989). Although the absolute values of the rate constants obtained from electrochemical kinetics are not precise (since the true surface concentration of peroxidase remains unknown) this method can be used to compare various forms of recHRP as a first approach, assuming their equal concentration on the electrode surface.

## Results and Discussion

### Light Scattering

The results of dynamic light scattering are represented as particles size distributions obtained by CONTIN program. In the case of recHRP (Fig. 2a) an asymmetric size distribution supposes a multimodal molecular motion. Cumulant analysis of the shape of the time correlation function allows us to suggest that there have three modes of molecular motion in the solution, corresponding to the sizes of monomeric ( $\sim 3.2\text{ nm}$ ), dimeric ( $\sim 6.4\text{ nm}$ ), and possibly tetrameric recHRP (Table 1). The native HRP in the same buffer (Fig. 2b) exhibits a size distribution different from that of

recHRP. Large characteristic sizes of the scattered particles (Table 1) and significant polydispersity (the width of distribution) testify that the native enzyme forms aggregates. The addition of salt crucially changes the size distribution for native HRP (Fig. 2c) shifting it to smaller sizes similar to those observed in the case of recHRP (cf. Fig. 2a). The value of this z-average hydrodynamic diameter is ca. twice larger than that of the monomer. The higher hydrodynamic radius can be attributed to the oligosaccharide “fur” surrounding the native enzyme.

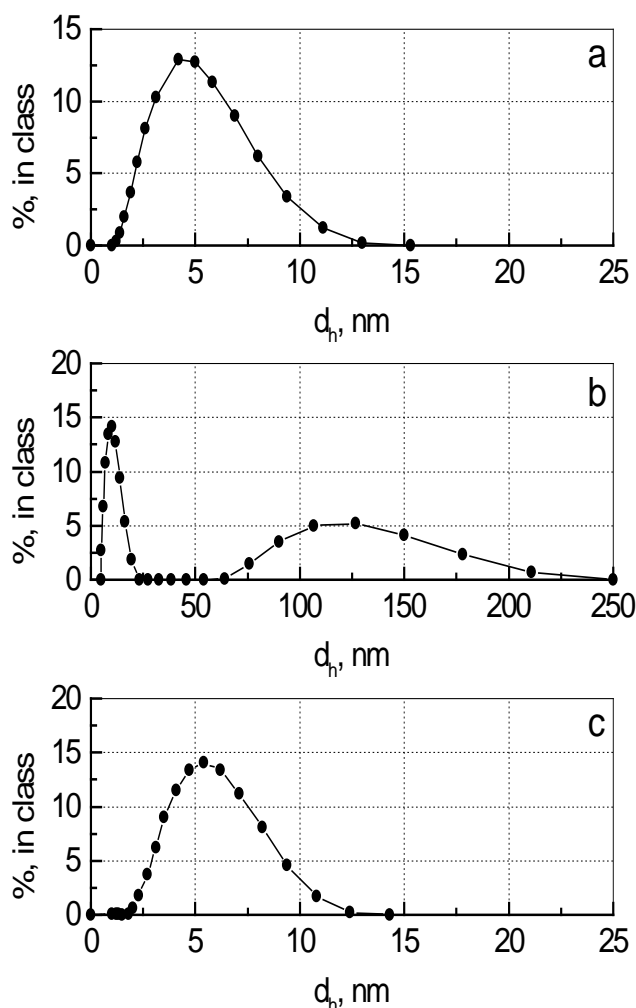


FIG. 2 PARTICLE SIZE DISTRIBUTIONS FOR ENZYMES SOLUTIONS IN ACETIC BUFFER AT pH 8: a) recHRP, b) NATIVE HRP, c) NATIVE HRP CONTAINING 0.5 M NaCl

TABLE 1 CHARACTERISTIC SIZES (Z-AVERAGE HYDRODYNAMIC DIAMETERS) OF SCATTERED PARTICLES FOR THE SOLUTIONS OF RECOMBINANT AND NATIVE HRP

Enzyme	RecHRP	Native HRP	Native HRP 0.5 M NaCl
$d_H$ (Mode 1), nm	$3.2 \pm 0.2$	-	-
$d_H$ (Mode 2), nm	$6.4 \pm 0.5$	-	$6.7 \pm 0.5$
$d_H$ (Mode 3), nm	$13.1 \pm 1.0$	$13.3 \pm 1.0$	$9.1 \pm 1.0$
$d_H$ (Mode 4), nm	-	$167 \pm 8.0$	-

The light scattering data unequivocally prove that the

oligomeric states of recombinant and native enzyme in solution are strikingly different. Presumably, it is this difference that has effect on the integral characteristics of the enzyme such as stability in solution and in the course of physical immobilization on surfaces.

### Cross-Linking

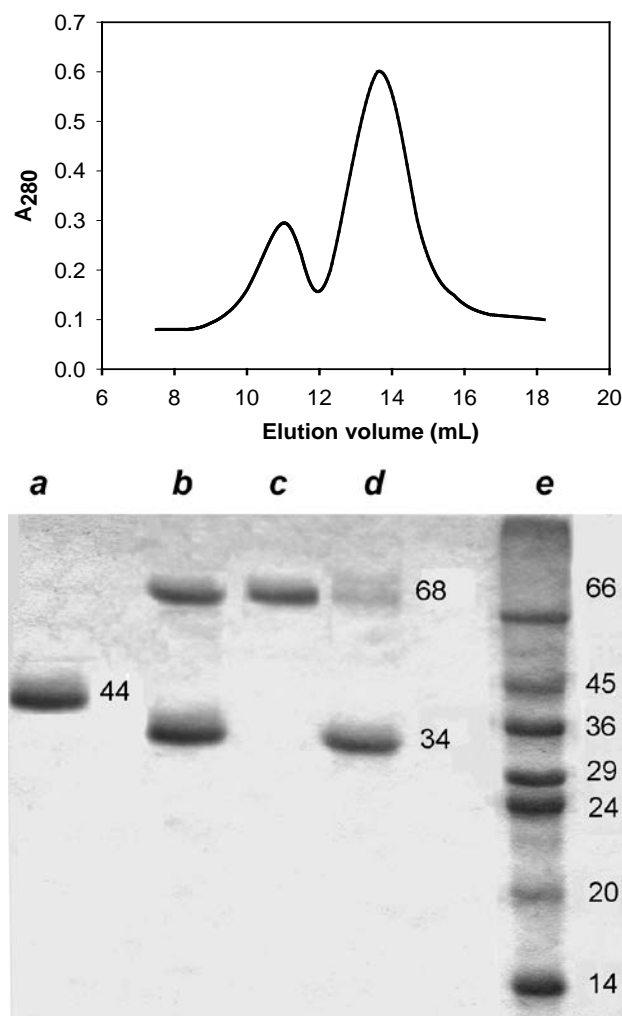


FIG. 3 GEL-FILTRATION PROFILE (UPPER) AND SDS-PAGE OF HRP PREPARATIONS AFTER CROSS-LINKING AND SEPARATION (LOWER): (a) NATIVE HRP AFTER A CROSS-LINKING ATTEMPT; (b) THE PRODUCT OF recHRP CROSS-LINKING; (c) DIMERIC, AND (d) MONOMERIC recHRP PREPARATIONS AFTER GEL-FILTRATION; (E)  $M_w$  MARKERS

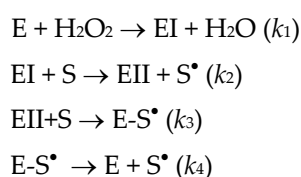
An attempt to obtain dimeric recHRP by cross-linking in the system of reversed micelles yields no more than 10% of dimeric enzyme (results not shown). In contrast, in water solutions, the very first attempt resulted in a 40% yield of the dimeric form, which was easily separated from the monomeric one by gel-filtration (Fig. 3). Of note, both enzyme preparations were separated from unreacted glutaraldehyde by this procedure. The dimeric fraction was homogeneous, while the monomeric fraction still contained ca. 10% of the dimer (Fig. 3). The cross-linking of native HRP (Fig. 3)

was inefficient: no enzyme dimers or aggregates were observed. Glutaraldehyde is a rather small linker and requires two a.a. residues to be at a close distance (6-7 angstrom) to be cross-linked. Heavily glycosylated and aggregated nHRP does not expose candidate a.a. residues on the surface, and shows no cross-linking products: there is a single band on SDS-PAGE (Fig.3). Contrary to nHRP, recHRP has Lys 232 residues sufficiently close in a dimer structural model, and these lysines can be considered as potential candidates for intramolecular cross-linking in recHRP dimer (Fig. 4). It is clear from this structural model that glycosylated Asn 268 and Asn198 in nHRP will interfere with the formation of the optimized dimeric structure, and thus, modification of nHRP lysines with glutaraldehyde does not yield any cross-linked oligomeric structure, despite that it is known to increase the enzyme stability (Ugarova et al., 1979).

### Changes in Substrate Specificity

The specific activity of monomeric and dimeric recHRP compared to the original preparation of the wild-type recHRP was changed as shown in Fig. 5. In all cases modification with glutaraldehyde causes a drop in the enzyme specific activity. This drop in activity is similar to monomeric and dimeric enzyme in the case of phenolic substrates, and may reflect the fact that extensive modification of the enzyme surface with glutaraldehyde interferes with the access of peroxidase substrates to the active site. In the case of organic amines such as ABTS, o-PD, and o-DA, the activity of the dimeric enzyme is ca. 50% of that of the monomeric (Fig. 5). The distinct effects in the case of phenolic and amine substrates may indicate principal differences in the mechanism of their oxidation. To evaluate the effects of cross-linking on the individual steps of the peroxidase catalytic cycle, we have studied ABTS oxidation (Table 2). The reaction of ABTS oxidation with hydrogen peroxide catalyzed by wild-type recHRP includes a rate-limiting monomolecular step, which is ascribed to the product dissociation step by Smith et al. (Smith et al., 1992) as shown in Scheme 3.

#### SCHEME 3



As evident from the values of the rate constants presented in Table 2, cross-linking has almost no effect on the properties of the monomeric enzyme, while it has

effect on the enzyme reduction steps in the case of the dimer, in particular, by decreasing the monomolecular rate-limiting constant  $k_4$  corresponding to the product dissociation step. It can be speculated that cross-linking results in a dimeric structure, where the active sites are almost facing each other (Fig. 4) and therefore, interaction with the oxidation product may interfere with the access of the substrate to the adjacent subunit. Alternatively, it is inferred that all amine substrates are oxidized via electron transfer through the protein molecule, and therefore, the decrease in the available surface per active center in the dimer results in the lower rate constant for the substrate oxidation step.

TABLE 2 RATE CONSTANTS FOR ABTS OXIDATION BY DIFFERENT FORMS OF RECHRP

Enzyme form	$k_1/10^7, \text{M}^{-1}\text{s}^{-1}$	$k_3/10^6, \text{M}^{-1}\text{s}^{-1}$	$k_4/10^3, \text{s}^{-1}$
Wild-type recHRP	1.2±0.2	2.5±0.5	2.6±0.5
Monomeric recHRP	1.0±0.2	1.6±0.3	2.4±0.5
Dimeric recHRP	1.0±0.2	0.8±0.2	1.3±0.3

As it was previously reported (Gazaryan et al., 1997) the presence of a particular substrate in the experiments with reversed micelles shifted the equilibrium between monomeric and dimeric recHRP forms. The most pronounced effect was observed for o-DA, which decreased the amount of the dimeric form up to 2-fold. A lesser effect was observed for pyrogallol, where the amount of the dimer was decreased by 20% (Gazaryan et al., 1997). Guaiacol had no effect on the equilibrium at all. These observations are in agreement with the data obtained in this work and support the first explanation, i.e. steric hindrance for the access of bulky substrates to the active site in recHRP dimer.

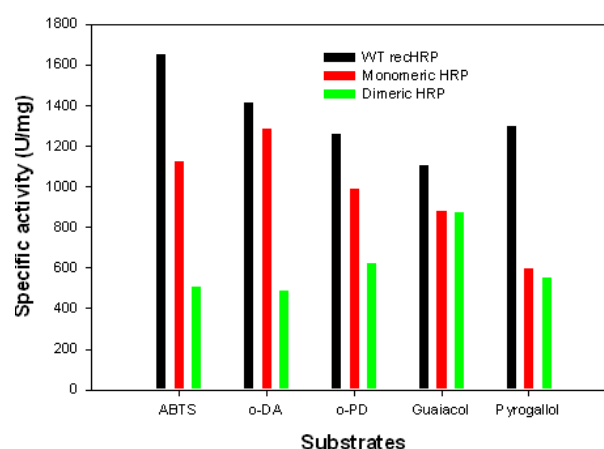


FIG. 5 SUBSTRATE SPECIFICITY OF recHRP, AND ITS MONOMERIC AND DIMERIC FORMS OBTAINED IN THE COURSE OF CROSS-LINKING WITH GLUTARALDEHYDE

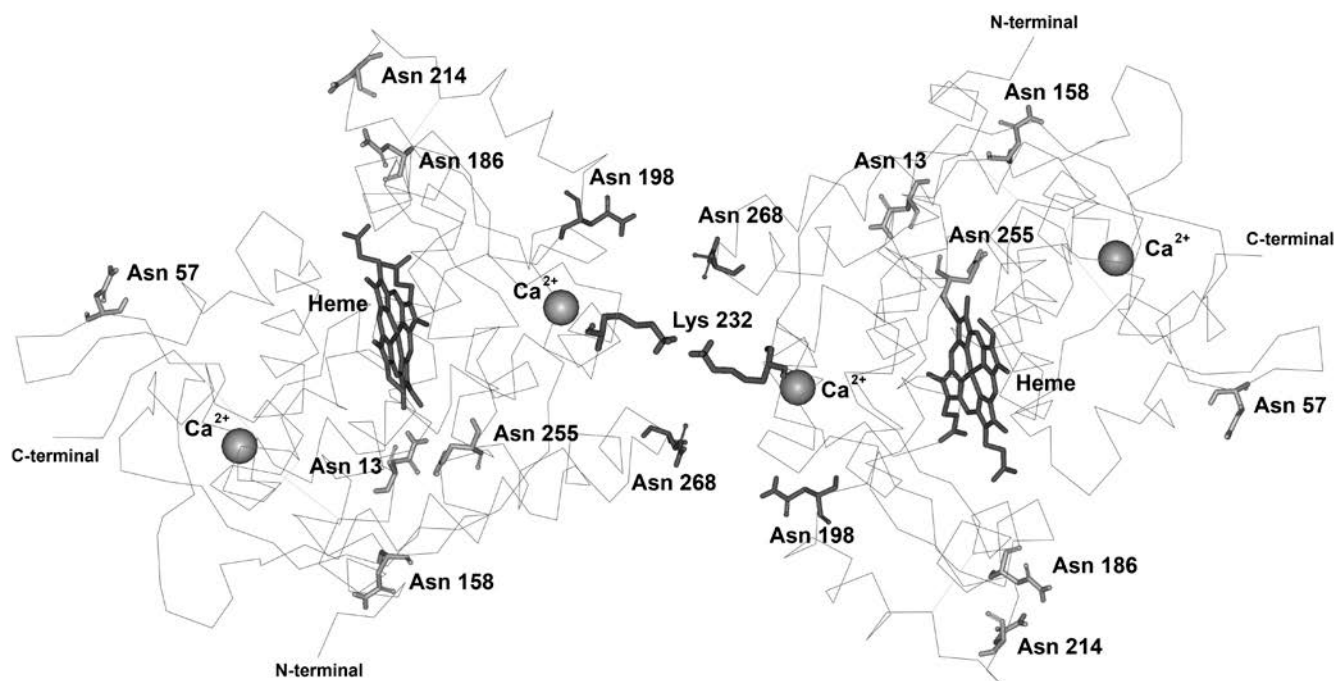


FIG. 4 COMPUTER MODEL OF DIMERIC *recHRP*: ASN RESIDUES GLYCOSYLATED IN NATIVE PEROXIDASE ARE SHOWN IN GREY TO DEMONSTRATE THAT OLIGOSACCHARIDE CHAINS ATTACHED TO ASN 268 AND ASN 198 WILL INTERFERE WITH ENZYME DIMERIZATION AND INTRAMOLECULAR CROSS-LINKING OF LYS 232

On the other hand, *o*-PD was shown to increase the amount of dimeric recHRP by 20% in the reversed micelles (Gazaryan et al., 1997), while in this work the substrate specificity of the dimer towards *o*-PD exhibits a 30% drop compared to the monomeric enzyme. The distinct results obtained with *o*-PD for the ferric form of recHRP in reversed micelles and for recHRP dimer in the reaction course may reflect the high affinity of the enzyme for both substrate and oxidation product, and thus, the decreased rate constant for the product dissociation, as in the case of ABTS. Overall results obtained in this work unequivocally demonstrate that cross-linking imposes more restrictions for substrate access to the active site in dimeric recHRP compared to the modified monomeric enzyme and especially unmodified recHRP.

### Electrochemical Kinetics

The method developed by us earlier (Gorton et al., 1999; Lindgren et al., 1998; Ruzgas et al., 1995) was applied to estimate the kinetic parameters for the monomeric and dimeric recHRP immobilized on graphite electrodes and compared with previously reported values for the unmodified recHRP and native enzyme (Table 3). First, the amount of enzyme available for direct electron transfer is 2 times lower in the case of the dimeric enzyme than that in the case of the monomeric one. This observation can be rationalized if one imagines that the dimer actually forms a kind of a

“double” layer as shown in Fig. 6B. In this case, the actual concentration of active centers on the electrode surface is doubled compared to that for the monomeric enzyme, and so, we must divide the kinetic parameters by 2 to account for the double concentration of active sites in the case of dimeric recHRP, leading to the conclusion on equal rate constants for all catalytic steps for the monomeric and dimeric enzyme within the experimental error. This conclusion is in agreement with the substrate specificity data for both enzyme forms in the case of phenolic substrates (it was reminded that *p*-cresol, a phenolic substrate, was used in the electrochemical experiments). Second, a comparison of the amount of enzyme in DET for the modified monomeric and unmodified recHRP with that for native HRP shows that modification with glutaraldehyde has the same effect as the presence of oligosaccharide chains in the native enzyme, i.e. screens a significant portion of enzyme surface and/or active sites available for direct electron transfer. The dramatic improvement in all kinetic constants observed for the recHRP forms compared to the native enzyme has now to be interpreted as the enzyme structural stabilization providing maximum activity yield in the course of physical immobilization on the graphite surface. The hydrophilic native enzyme, heavily aggregated in water solutions, undergoes major inactivation in the course of interaction with graphite surface. This is the main reason for the unbelievably low, at least 2 orders of magnitude, values for the rate constants of the indi-

vidual steps of the peroxidase catalytic cycle estimated from electrochemical experiments for the native enzyme (Table 3) compared to those in water solutions (Table 2). The recombinant monomeric HRP shows almost an order of magnitude increase in all rate constants compared to the native enzyme (Table 3), and a 2-fold increase compared to the unmodified recHRP. Thus, additional hydrophobization of the recHRP surface has an extremely positive effect in terms of minimizing the enzyme inactivation in the course of its immobilization on electrodes. However, the achieved values for  $k_1$  and  $k_3$  are still lower than those in solution by 20- and 5-fold, respectively.

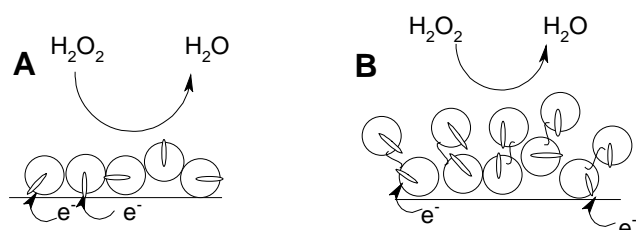


FIG. 6 SCHEMATIC REPRESENTATION OF "DOUBLED" CONCENTRATION OF ACTIVE SITES IN THE CASE OF EQUAL MOLECULAR COVERAGE OF THE ELECTRODE SURFACE IN THE CASE OF MONOMERIC (A) AND DIMERIC (B) RECHRP

This difference actually points to the fact that the  $\text{H}_2\text{O}_2$  cleavage site and the substrate binding site are structurally different because they exhibit different stability with respect to immobilization by physical absorption. It is concluded that the peroxidase active site featuring a unique disposition of catalytic histidine and arginine toward the heme iron is much more sensitive to structural damage than the substrate binding sites in the peroxidase molecule.

TABLE 3 KINETIC CHARACTERISTICS OF DIFFERENT FORMS OF HRP IMMOBILIZED ON GRAPHITE ELECTRODE. AVERAGE OF 3 ELECTRODES, 95% CONFIDENCE LIMITS GIVEN. 0.1 M NA-PHOSPHATE, PH 7.0. THE RATE CONSTANTS ARE CALCULATED USING A MONOLAYER COVERAGE (NUMBER OF HRP ACTIVE SITES/AREA)

	Native HRP (Lindgren et al., 1998)	RecHRP (Lindgren et al., 1999)	Mono- meric recHRP	Dimeric recHRP
% in DET	48.0±4.0	63.0±7.0	42.0±5.0	19.0±5.0
$k_1/10^5$ $\text{M}^{-1}\text{s}^{-1}$	1.3±0.2	2.8±0.6	6.0±3.1	11.6±2.5
$k_s$ $\text{s}^{-1}$	1.9±0.3	7.6±2.5	13.0±6.0	21.0±9.0
$k_3/10^5$ $\text{M}^{-1}\text{s}^{-1}$	0.3±0.1	1.4±0.7	2.3±1.4	5.0±4.0
Number of electrons	2.1±0.2	1.8±0.2	1.94±0.06	1.90±0.06

In conclusion, the revealed tendency of the native glycosylated peroxidase to form aggregates in water solu-

tions is of key importance for all protocols based on physical adsorption of the enzyme on surfaces, and in particular immobilization of electrodes. Immobilization by adsorption is known to aggregate and inactivate proteins (Kondo et al., 1993), and obviously, for the protein already aggregated in solution, the inactivation effect in the course of immobilization will be even stronger. It is interesting to note that the recent study aimed at production of HRP encapsulated in hollow gold nanoparticles was successful due to the use of HRP micelles which provided the entrapment of individual molecules first into the micelle and then used as a basic structure to make a 100 nm nanoparticle (Kumar et al., 2005).

The results obtained in this study allow us to demonstrate that formation of dimeric structures by the recombinant wild-type HRP is the inherent property of the enzyme molecule, which is preserved both in water solutions and in micellar systems. The demonstrated increase in operational characteristics of immobilized recHRP by chemical cross-linking in water solutions can be used when studying enzyme applications to biosensors and in enzyme immunoassay. The major practical conclusion from our studies is that for immobilization on an electrode one needs monomeric HRP (native or recombinant), no matter how this will be achieved, i.e. reversed micelles, cross-linking and separation for recHRP, high salt for nHRP, or entrapment into hollow nanospheres (Xie et al., 2011).

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